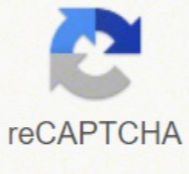




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Polyatomic Practice

- Name or write the formula for the following polyatomic ions

| | | | |
|--------------|-------|---------------------|-------|
| sulfate | _____ | CO_3^{2-} | _____ |
| nitrite | _____ | MnO_3^{1-} | _____ |
| perphosphate | _____ | SO_5^{2-} | _____ |
| hypoiodite | _____ | BrO_2^{1-} | _____ |
| chlorite | _____ | CO_4^{2-} | _____ |
| phosphite | _____ | PO_3^{3-} | _____ |
| percarbonate | _____ | ClO^{1-} | _____ |
| bromate | _____ | IO_2^{1-} | _____ |
| hyposulfite | _____ | PO_4^{2-} | _____ |
| permanganate | _____ | NO_3^{1-} | _____ |
| carbonite | _____ | SO_4^{2-} | _____ |

- Name or write the formula for the following Type I polyatomic ionic compounds

| | | | |
|------------------------|-------|---|-------|
| beryllium hydroxide | _____ | $\text{Ba}(\text{IO}_3)_2$ | _____ |
| sodium nitrite | _____ | $\text{Ga}(\text{CNO})_3$ | _____ |
| ammonium chloride | _____ | Ag_2SO_3 | _____ |
| calcium bisulfate | _____ | MgCO | _____ |
| rubidium perchlorate | _____ | NH_3NO_2 | _____ |
| strontium sulfite | _____ | $\text{Al}(\text{C}_2\text{H}_3\text{O}_2)_3$ | _____ |
| aluminum acetate | _____ | SrSO_3 | _____ |
| ammonium nitrate | _____ | PbClO_2 | _____ |
| magnesium hypcarbonite | _____ | $\text{Ca}(\text{HSO}_4)_2$ | _____ |
| silver hyposulfite | _____ | NH_4Cl | _____ |
| gallium cyanate | _____ | NaNO | _____ |
| barium iodate | _____ | $\text{Be}(\text{OH})_2$ | _____ |

Identifying Covalent Binary Compounds: Elemental Molecular Compounds

Write the name of each of the following compounds:

- | | | |
|---------------------------|--|-------------------------|
| 1. N_2 | | 1. nitrogen monoxide |
| 2. N_2O | | 2. dinitrogen monoxide |
| 3. NO | | 3. nitrogen monoxide |
| 4. N_2O_2 | | 4. dinitrogen dioxide |
| 5. NO_2 | | 5. nitrogen dioxide |
| 6. N_2O_5 | | 6. dinitrogen pentoxide |
| 7. SO_2 | | 7. sulfur dioxide |
| 8. SO_3 | | 8. sulfur trioxide |
| 9. CO_2 | | 9. carbon dioxide |
| 10. CO | | 10. carbon monoxide |

Write the chemical formula for each of the compounds:

- | | | |
|--------------------------|----------|----------------------------|
| 11. nitrogen dioxide | | 11. NO_2 |
| 12. dinitrogen monoxide | | 12. N_2O |
| 13. sulfur trioxide | | 13. SO_3 |
| 14. carbon monoxide | | 14. CO |
| 15. dinitrogen pentoxide | | 15. N_2O_5 |
| 16. nitrogen monoxide | | 16. NO |
| 17. carbon dioxide | | 17. CO_2 |
| 18. nitrogen monoxide | nitrogen | 18. NO |
| 19. carbon monoxide | carbon | 19. CO |
| 20. sulfur dioxide | sulfur | 20. SO_2 |

Variable charge cations worksheet answer key.

For example, you could see the words stagnant fluoride on a toothpaste tube. The simplest name, FeCl_2 , "Iron chloride," in this case, be ambiguous, as it does not distinguish between these two compounds. The charge of the metal ion is determined by the formula of the compound and from the charge of the anion. These accusations are used in the names of the metal ions: iron (iii) copper sulfide (ii) selenide gallium (iii) nitride chromium (iii) chloride titanium (iii) sulfate exercise (PageDex { 1 }) Response Fluoro: (a) CRP; (b) HGS; (C) MN3 (PO4) 2; (D) CU2O; (E) CRF6 For example, considering iron and chlorine binary ion compounds. Example (PageIDyx { 1 }): Name ionic compounds Name of the following ionic compounds, which contain a metal that may have more than an ionic charge: FE2S3 CUSE GAN CRCL3 TI2 (SO4) 3 SOLUTION The anions in these compounds have solved uploads Tiva (S2A, SE2A, N3A, cIA, and (sot4 ^ 2 -)), and the compounds must be neutral. Iron typically shows a 2+ or 3+ charge (see [Connection]) and the two corresponding composed formulas are FECL2 and FECL3. The other fluoride of Tin is SNF4, which has previously been called stannic fluoride, but now it is called Tin Fluoride (IV). In cases like this, the metal ion charge is included as a Roman number in brackets immediately following the name of the metal. Most transition metals can form two or more cations with different charges. Other examples are provided in the table. The compounds of these metals with non-metals are appointed with the same method of the compounds in the first category, except for the metal ion charge it is specified by a Roman number in brackets after the III(III niT atanionned etnematterroc 'Aip 'A ehc .2FNS alumrof al atneserppar otseuQ .ollatem led yrtsudni .yrtsudni fo stnenges emos yb esu ni sniamer ti, ytiammoc ciftneies eht yb donodnaba ylegral neb donodnaba ylegral neb habuote yliman siht hguoht. + 3it dna, + 4rc, + 3aq, + 2uc, + 3ef eh tsum snoi evitisop eht, segrahc ivitagen fo Rehmun latot eht lauge tsum dnuopmoc hcae ni sulfrahc evitisop fo er esuaceb .edirohlc Suorref SA nwonk Saw, 2lcef, edirohlc) 2lcef, edirohlc cirref dellac ylsuoiverp saw, 3lcef, edirohlc) 3LCEF, edirohlc) 3lcef, ylevicepser, segrahc rewol dna Rehghit ot sua "a C a dna cia" a C e e sexiffus eht desu erutalcnemom ehti erutalcnemom etad-fo-your EtahpsoHP) 2 (Reppoco 2) 2 (3rem Ogh edixo) 1 (yrucrem OZGH edirohlc) 3 (Nori 3LCEF eman DnuopMoc Cinoi Late Noittisart SDNUOPMOC CINOI NORIONART emos FO Seman:) \} 3 {xedniegap \ (elibat .Levitcepser, edirohlc) 3 (Nori DNA edirohlc) 2 (Nori Deman ylsuougibmanu neht era sduopmoc owt eseht

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