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All rights reserved AnswerVerifiedHint: We know that the coenzyme helps the enzyme to do its process faster to increase the rate of reaction by converting reactants into products faster and more easily by lowering the activation energy of the overall reaction.Equilibrium constant is defined as the constant when the rate of forward reaction is equal to the rate of reverse reaction. The concentration of reactants and products are constant at equilibrium. The value of equilibrium constant can be determined if we know the concentration of the reactants and products.Complete answerThe reaction will start taking place in the presence of enzyme and coenzyme. It will help to break the bond of the reactants molecules by providing the proper surface and the required reaction conditions. The enzyme will act as a lock in which reactants will act as key buy inserting it in the enzyme it will modify the reactants into the desired products and coenzyme will help enzyme to increase the rate of reaction in forward direction.The product will be detached and again we will get the enzyme which is reusable. The product obtained will be extracted. During this process the desired catalyst or enzyme used will lower the activation energy hence the reactants will easily get converted into the product and the concentration will remain the same. The condition will come at which the concentration of reactants is equal to that of product, it is an equilibrium condition, the value of that is equilibrium constant and it is dependent on concentration of reactants and not on the catalyst.Hence, the incorrect option is (C). The value of equilibrium constant is changed in the presence of catalyst in the reaction at equilibriumNote: The equilibrium constant can be used to determine the half life time of the reaction. It gives us information of the concentration of the reactants used up and the product formed at different time interval.it also helps us to find the age of the old objects which contain carbon atoms.AnswerVerifiedHint: Catalyst is a common word that you might come across while studying chemistry especially while learning about chemical reactions. While some of the chemical reactions occur quickly, some take a long time and require extra materials or effort. So, catalysts are defined as those substances which alter the rate of reaction by changing the path of reaction.Complete answer:A catalyst can be either solid, liquid or gaseous catalyst. A catalyst increases reaction rates in a slightly different way from other methods of increasing reaction rate. The function of a catalyst is to lower the activation energy so that a greater proportion of the particles have enough energy to react.Types of Catalysts are:1) Positive Catalysts : Catalysts which increase the rate of a chemical reaction are positive catalysts.2) Negative Catalysts: Catalysts which decrease the rate of reaction and negative catalyst.The function of a catalyst is to lower the activation energy so that a greater proportion of the particles have enough energy to react.A catalyst increases the rate of reaction by decreasing the activation energy of the reaction. It does not affect the Gibbs free energy \[\Delta G\] of the reaction. Thus, the statement given in option (B) is incorrect.Note: In essence, catalysts encourage molecules to react and make the whole reaction process easier and efficient.AnswerVerified Hint: Catalysts are the substances which increase the rate of a reaction under different conditions. Catalysts are not consumed in a reaction.Complete Step by Step Solution:A. When a catalyst is added in a reaction, it does not take part in the reaction but affects the reaction and the quantity of product at the end of the reaction. Therefore, the statement is true.B. The catalyst increases the rate of reaction but does not initiate a reaction. Therefore, the statement is false.C. A catalyst pushes the reaction in both the forward and backward reaction in a reversible reaction. Therefore, the statement is true.D. Catalysts are specific for different types of reactions. Different reactions use different catalysts. Therefore, the statement is true.Therefore, options A, C and D are correct.Additional information: The process by which the reaction rate is increased by adding a catalyst is known as catalysis. A few examples of catalysts are palladium, platinum, iron, nickel, and zeolites.Note: Catalyst and enzyme have the same effect on the reaction but they are different from each other. Enzymes are naturally occurring whereas catalysts can be synthesised. All enzymes act as a catalyst but all catalyst is not an enzyme.It alters G of the reaction.Explanation:(i) It catalyses the forward and backward reactions to the same extent as it decreases energy of activation hence, increases the rate of both the reactions.(ii) Since reaction quotient is the relation between concentration of reactants and products. Hence, catalyst does not alter Gibbs free energy as it is related reaction quotient. Thus, Gibbs free energy does not change during the reaction when catalyst is added to it.(iii) It does not alter equilibrium of reaction as equilibrium constant is also concentration-dependent term.(iv) It provides an alternate mechanism by reducing activation energy between reactants and product.->NEET 2025 exam is over on May 4, 2025. -> Check the LIVE NEET Exam Analysis 2025 here. -> The NEET 2025 Question Papers PDF are now available. -> NEET Admit Card 2025 is out on April 30, 2025. -> NTA has released the NEET City Intimation Slip 2025at on April 23, 2025. ->The applications for NEET (UG) were accepted from7th February 2025 to 7th March 2025. ->NTA has changed the NEET UG Exam Pattern of the NEET UG 2025. Now,there will be no Section B in the examination. -> Candidates preparing for the NEETExam, can opt for the latest NEET Mock Test 2025. -> NEET aspirants can check the NEET Previous Year Papers fortheir efficient preparation. and Check NEET Cut Offhere.

Which one of the following statement is not correct catalyst does not initiate any reaction.